

Ionic Liquid

Subjects: Chemistry, Physical

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Ionic Liquids (ILs) are organic salts that melt commonly below 100 °C, constituted entirely by charged species. The tunability and versatility of ILs have given rise to several applications at the academic and industrial levels. Here the following topics are highlighted: chemical structure; ILs classification according to the possibility of proton transfer, and the historical generations of ILs.

Keywords: organic salts ; tunable properties ; intermolecular interaction

1. Introduction

The extensive amount of possible combinations of known cations and anions (10^6 – 10^{18})^[1], results in different and unique physicochemical properties such as high thermal stability, large electrochemical window, and low vapor pressure. Because of these attractive properties, ILs have garnered industrial and scientific interest, see Figure 1^{[2][3]}.

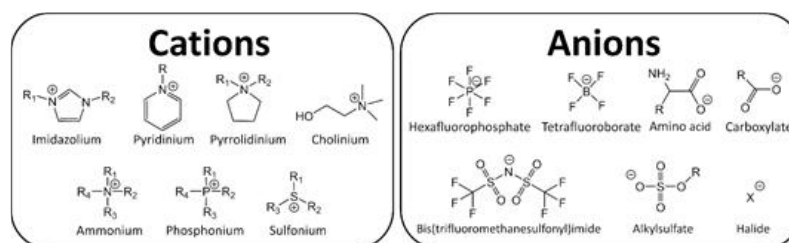


Figure 1. Typical ions in ionic liquids (ILs).

Imidazolium-based ILs (ImILs) is one of the most popular classes of organic salt explored in the literature. They can be described as a supramolecular polymeric network of intermolecular interactions that determine the interionic distances and consequently the existence of ion pairs and/or aggregates and free volume. The local nanostructural organization and the physicochemical properties of ILs are directly related. The interrelation between transport properties (viscosity, diffusion, conductivity) of IL structure and free volume was reviewed (see link below), analysing the microscopic features to explain and predict macroscopic properties, reaching new perspectives on the properties and application of ILs.

The tunability and versatility of ILs have given rise to several applications such as solvents for synthesis and catalysis^{[4][5]} CO₂ capture and storage^{[6][7][8]} energy generation and storage^[9], extraction/dissolution of biomass^{[10][11]}, and active pharmaceutical ingredients^{[12][13]}.

2. Classification

ILs are normally divided into two classes: aprotic ionic liquids (APILs) and protic ionic liquids (PILs). PILs are formed by the transfer of protons from Brønsted acid to Brønsted base, this proton transfer does not take place in APILs due to the nature of the ions constituting the salt. PILs are generally prepared through a neutralization reaction and APILs by a quaternization reaction followed by anion exchange, Figure 2^{[14][15]}.

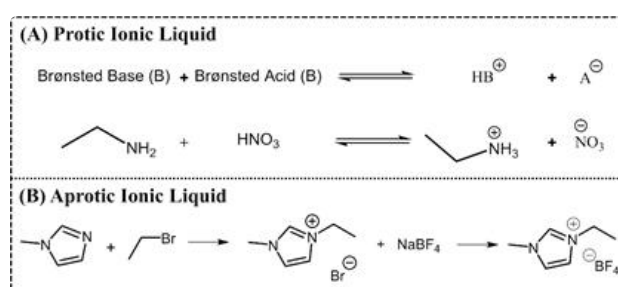
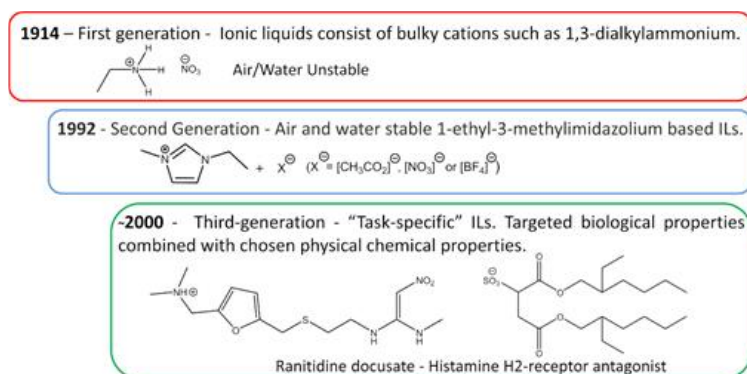


Figure 2. Typical synthetic route of a protic ionic liquid (PIL) (A) ethylammonium nitrate and an APIL (B) 1-ethyl-3-methylimidazolium tetrafluoroborate.

3. Different Generation of ILs

The first IL reported in the literature is a controversial matter. Most authors state that the history starts with Paul Walden's discovery in 1914 with the report of ethylammonium nitrate ($[\text{EtNH}_3][\text{NO}_3]$) synthesis, see Figure 3^[16]; however, some authors attribute the first IL to Gabriel and Weiner in 1888^[17], with the synthesis of ethanolammonium nitrate. This confusion arises from the reported melting point, while Walden's compound was characterized with a melting point of 13–14 °C, Gabriel's compound exhibited a melting point of 50 °C, substantially higher than room temperature. These works marked the beginning of the First Generation of ILs. Cations of the first generation of ILs are characterized by large volumes, such as 1,3-dialkyl-imidazolium or 1-alkylpyridinium and anions based mostly on halogen aluminate (Al^{3+})^{[1][18]}. The disadvantage of this generation was the instability in relation to air and water.



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