

AOPs for Water Treatment

Subjects: [Water Resources](#)

Contributor: Joaquim C.G. Esteves Da Silva

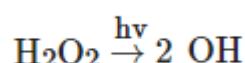
Advanced oxidation processes (AOPs) are water treatment processes that are promising for the degradation of persistent or toxic organic pollutants, as well as compounds refractory to other environmental remediation/decontamination treatments. AOPs have gained great importance as alternative treatment processes that affect the degradation of organic species through the action of the hydroxyl radical (OH), oxidizing pollutants present in wastewater and industrial effluents. AOPs are carried out at room temperature and at a pressure close to normal, which involve the formation of very reactive radical species with a high oxidizing capacity, mainly hydroxyl (OH) radicals. These OH radicals are extremely reactive oxidizers (oxidation potential of the OH radical is approximately, $E\theta = 2.8$ V) and non-selective towards organic pollutants in wastewater. AOPs can be considered versatile technologies, as they provide different possible alternatives to produce OH radicals. AOPs, compared to conventional water treatment techniques, have a greater efficiency and capacity to degrade recalcitrant organic pollutants, and can generate less toxic intermediate products during their degradation.

AOPs

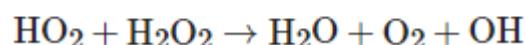
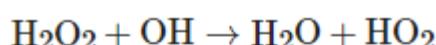
water treatment

1. UV/H₂O₂ Processes

The AOP based on the UV/H₂O₂ system consists of the combination of ultraviolet radiation with hydrogen peroxide (H₂O₂), leading to its photolysis, in order to generate two hydroxyl (OH) radicals ^[1], as illustrated in following equation ^[2]:



In addition to direct photolysis, the UV/H₂O₂ process can also comprise indirect photolysis, where the compounds react with the OH radical, produced by the photolysis of H₂O₂, in order to generate less reactive radicals such as HO₂, which can react with the H₂O₂, generating again the OH radical ^{[1][3]}, as shown in the following equations:



The efficiency of OH radical production depends on the ability of hydrogen peroxide to absorb UV radiation, as well as the physical and chemical characteristics of the fluid that will be subjected to the oxidation process. UV absorption by H_2O_2 will depend on the wavelength of UV radiation, that is, the shorter the wavelength of UV radiation, the greater the energy absorption by H_2O_2 , increasing the production potential of the OH radical [2].

An advantage of the UV/ H_2O_2 system is due to the fact that UV radiation acts as a disinfectant, physically inactivating the microorganisms and simultaneously assisting the peroxide photolysis, generating highly reactive OH radical species. Thus, the UV/ H_2O_2 system is currently one of the most promising technologies for wastewater treatment [2][4].

2. Persulfate Based Processes

Recently developed, the persulfate oxidation technology ($\text{S}_2\text{O}_8^{2-}$) is a new AOP, which has been shown to be a very promising alternative for water and wastewater treatment [5][6].

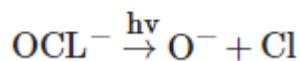
Due to its high oxidizing power ($E^\theta = 2.01$ V), persulfate is considered an emerging oxidant in the degradation of pollutants present in water, moreover, persulfate is relatively stable at room temperature and has a non-selective behavior in regarding the degradation of pollutants [7][8]. Persulfate can also be activated to produce the sulfate radical (SO_4), an even stronger oxidant ($E^\theta = 2.6$ V) than $\text{S}_2\text{O}_8^{2-}$ [9].

There are several methods of activating persulfate, in general, the SO_4 radical can be generated through heat, ultraviolet light, alkali, ultrasound, transition metal ions and activation of metal oxides [6][10].

Compared to the OH radical, the SO_4 radical has a longer shelf life, has a wider pH range of action, proved to be more stable and effective in oxidation for water decontamination [11][12]. Furthermore, some studies suggest that after the generation of the SO_4 radical, it may react with several species in solution to form other active species, such as the OH radical, which will play a key role in the pollutant degradation process [9][12].

3. Chlorine and NH_2Cl Based Processes

Chlorine is one of the most used chemical oxidants worldwide for the disinfection of drinking water [13]. Currently, the UV/chlorine process is considered an emerging AOP, constituting an alternative to the UV/ H_2O_2 process in water treatment, effective in the degradation of a variety of persistent contaminants, such as desethylatrazine, sulfamethoxazole, carbamazepine, diclofenac, benzotriazole, tolyltriazole, iopamidol, 17 α -ethinylestradiol [14][15]. Compared to the UV/ H_2O_2 process, the UV/chlorine process demonstrated greater efficiency in the degradation of some micropollutants under slightly acidic conditions, such as trichloroethylene [16]. In the UV/chlorine process, reactive species such as the OH radical and the Cl radical are formed from photolysis of free chlorine (HOCl/OCl^-), as shown in the following equations [15]:



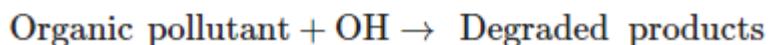
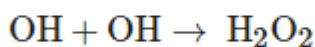
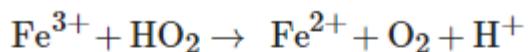
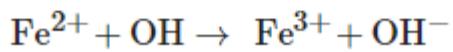
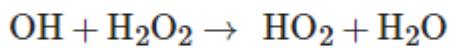
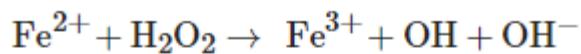
The Cl formed reacts with chloride, giving rise to $\text{Cl}_2^{\bullet-}$, and both Cl and $\text{Cl}_2^{\bullet-}$ are strong oxidizers with redox potentials of 2.4 and 2.0 V, respectively. The various reactive species, including the OH radical, formed during this process make the UV/chlorine process a promising AOP for controlling a variety of contaminants in water treatment [14]. Furthermore, chlorine that does not react after the UV/chlorine process can provide residual protection in water distribution systems [17].

Recently, photolysis of monochloramine (NH_2Cl) has also attracted significant interest as a new AOP for the degradation of emerging water contaminants, such as carbamazepine, and for efficiently controlling the formation of disinfection by-products [13]. Furthermore, NH_2Cl is considered adequate to provide residual disinfection throughout the water distribution system due to its high stability [18].

NH_2Cl photolysis generates NH_2 and Cl radicals, and due to their strong interaction with water, significant amounts of Cl radical are converted into OH radicals. As discussed previously the OH radical is a strong non-selective oxidizer against the different pollutants present in water. In contrast, Cl radical is a relatively selective oxidant that reacts with most organic micropollutants. In relation to the NH_2 radical, there is still not much information about its reactivity to pollutants present in water [19]. Primary radicals can further react with water co-solutes to form secondary radicals (such as $\text{CO}_3^{\bullet-}$ and $\text{Cl}_2^{\bullet-}$). Although knowledge is still quite limited, the UV/ NH_2Cl process demonstrates considerable potential as a new AOP for water treatment [13][19].

4. Fenton Processes

The Fenton process consists of the formation of OH radicals in an acidic medium, from the decomposition of hydrogen peroxide by the action of an iron catalyst [20]. The Fenton process is a very promising AOP due to the high mineralization promoted under normal conditions of both temperature and pressure, very effective in destroying refractory and toxic organic pollutants present in wastewater [21][22]. The main reactions involved in Fenton processes are [23]:



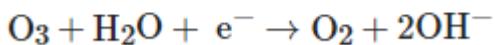
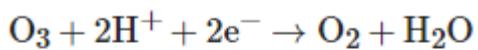
The advantage of Fenton processes is due to the fact that they do not require sophisticated equipment or expensive reagents, they are considered ecologically viable processes due to their high performance and their relatively simple approach, which uses less harmful chemicals and cyclical nature, being necessary a lower concentration of these chemicals [23].

A condition of this AOP is the restricted pH range, where the generation of OH radicals during the Fenton reaction is only efficient under acidic pH conditions (with a pH value close to 3) [21][22].

There are different types of Fenton processes, including Fenton, photo-Fenton, electro-Fenton, photo-electro-Fenton, homogeneous and heterogeneous Fenton, among others. Among the various AOPs, it was the Fenton and photo-Fenton processes that proved to be the most effective, energy efficient and least expensive methods for treating recalcitrant compounds, when used exclusively or in conjunction with other conventional and biological methods [23].

5. Ozone Based Processes

Ozone (O_3) is a strong oxidizer, which has been widely used in wastewater treatment and drinking water disinfection. Once dissolved in water, O_3 acts as an oxidant due to its high redox potential ($E^\theta = 2.07 \text{ V}$), leading to the destruction or degradation of organic pollutants, through different pathways, namely molecular ozone (direct) or through the hydroxyl radical (indirect) [24]. The main semi-reactions of ozone in water are [25]:



However, direct oxidation of O_3 , in addition to being a relatively slow reaction, is also more selective than indirect reactions [26]. In water, O_3 undergoes a series of reactions, decomposing into several oxidative species, including the OH radical, which is a stronger oxidant than the original molecular ozone, which reacts with organic and inorganic compounds in a non-selective way, with rates very high reaction rates. Thus, the high oxidative power of ozonation is partially due to the generation of OH radicals [25].

In the presence of other oxidants or irradiation, the yield of the OH radical can be significantly improved, namely through the addition of hydrogen peroxide ($\text{O}_3/\text{H}_2\text{O}_2$), use of UV irradiation (O_3/UV), or addition of metallic catalysts, increasing the efficiency of the treatment. These processes are called advanced oxidation processes (AOPs). Thus, ozonation and ozone-based AOPs are responsible for the destruction of many recalcitrant organic compounds in water and wastewater, including pharmaceuticals and personal care products, solvents, surfactants and pesticides [24][25].

6. Heterogeneous Photocalalytic Processes

Heterogeneous photocatalysis is an AOP that has been widely studied in the last two decades. The principle of heterogeneous photocatalysis is associated with the activation of a semiconductor by the action of light, when the semiconductor and the reagent are in different phases, photocatalytic reactions are classified as heterogeneous photocatalysis [27].

The substrate that absorbs light and acts as a catalyst for chemical reactions is known as a photocatalyst, and all photocatalysts are generally semiconductors. When a photocatalyst is exposed to light of the desired wavelength (sufficient energy), photoexcitation of an electron occurs, which is promoted from the valence band to the conduction band. The absorption of a photon with energy equal to or greater than the separation of the bands (band gap) of the catalyst, thus creates a gap in the valence band. The electron and the gap migrate to the surface of the photocatalytic semiconductor where they act, respectively, as a reducing agent and an oxidizing agent. On the catalyst surface, the adsorbed water molecules are oxidized by the holes, to OH radicals, which can subsequently oxidize the organic matter in solution, transforming it into non-toxic or less harmful products, such as CO_2 and water [28].

Among the various existing photocatalysts, TiO_2 is currently the most studied in heterogeneous photocatalysis processes, essentially due to its physical and chemical characteristics. TiO_2 has high chemical and thermal stability, is non-toxic, inexpensive and has a relatively high efficiency [27].

As a green, highly efficient and ecological technology, heterogeneous photocatalysis for the treatment of organic pollutants present in wastewater has been shown to be a promising technology to face future environmental challenges [29].

References

1. Boczkaj, G.; Fernandes, A. Wastewater treatment by means of advanced oxidation processes at basic pH conditions: A review. *Chem. Eng. J.* 2017, 320, 608–633.
2. Mierzwa, J.C.; Rodrigues, R.; Teixeira, A.C.S.C. UV-Hydrogen Peroxide Processes. In Advanced Oxidation Processes for Wastewater Treatment: Emerging Green Chemical Technology; Elsevier: Amsterdam, The Netherlands, 2018; pp. 13–48.
3. Wu, C.; Linden, K.G. Degradation and byproduct formation of parathion in aqueous solutions by UV and UV/H₂O₂ treatment. *Water Res.* 2008, 42, 4780–4790.
4. Zeng, P.; Zhang, Q.; Zhang, X.; Peng, T. Graphite oxide–TiO₂ nanocomposite and its efficient visible-light-driven photocatalytic hydrogen production. *J. Alloys Compd.* 2012, 516, 85–90.
5. Zhou, Y.; Shi, J.; Li, Y.; Long, T.; Zhu, X.; Sun, C.; Wang, Y. Degradation kinetics and mechanism of bis(2-chloroethyl) ether by electromagnetic induction electrodeless lamp activated persulfate. *Chemosphere* 2020, 261, 127709.
6. Yang, L.; Bai, X.; Shi, J.; Du, X.; Xu, L.; Jin, P. Quasi-full-visible-light absorption by D35-TiO₂/g-C₃N₄ for synergistic persulfate activation towards efficient photodegradation of micropollutants. *Appl. Catal. B Environ.* 2019, 256, 117759.
7. Li, J.; Ji, Q.; Lai, B.; Yuan, D. Degradation of p-nitrophenol by FeO/H₂O₂/persulfate system: Optimization, performance and mechanisms. *J. Taiwan Inst. Chem. Eng.* 2017, 80, 686–694.
8. Zhang, M.; Chen, X.; Zhou, H.; Murugananthan, M.; Zhang, Y. Degradation of p-nitrophenol by heat and metal ions co-activated persulfate. *Chem. Eng. J.* 2015, 264, 39–47.
9. Liu, S.; Lai, C.; Li, B.; Zhang, C.; Zhang, M.; Huang, D.; Qin, L.; Yi, H.; Liu, X.; Huang, F.; et al. Role of radical and non-radical pathway in activating persulfate for degradation of p-nitrophenol by sulfur-doped ordered mesoporous carbon. *Chem. Eng. J.* 2020, 384, 123304.
10. He, L.; Chen, H.; Wu, L.; Zhang, Z.; Ma, Y.; Zhu, J.; Liu, J.; Yan, X.; Li, H.; Yang, L. Synergistic heat/UV activated persulfate for the treatment of nanofiltration concentrated leachate. *Ecotoxicol. Environ. Saf.* 2021, 208, 111522.
11. Hu, L.; Wang, P.; Liu, G.; Zheng, Q.; Zhang, G. Catalytic degradation of p-nitrophenol by magnetically recoverable Fe₃O₄ as a persulfate activator under microwave irradiation. *Chemosphere* 2020, 240, 124977.

12. Castilla-Acevedo, S.F.; Betancourt-Buitrago, L.A.; Dionysiou, D.D.; Machuca-Martínez, F. Ultraviolet light-mediated activation of persulfate for the degradation of cobalt cyanocomplexes. *J. Hazard. Mater.* 2020, 392, 122389.

13. Bu, L.; Zhou, S.; Zhu, S.; Wu, Y.; Duan, X.; Shi, Z.; Dionysiou, D.D. Insight into carbamazepine degradation by UV/monochloramine: Reaction mechanism, oxidation products, and DBPs formation. *Water Res.* 2018, 146, 288–297.

14. Fang, J.; Fu, Y.; Shang, C. The roles of reactive species in micropollutant degradation in the UV/free chlorine system. *Environ. Sci. Technol.* 2014, 48, 1859–1868.

15. Kong, X.; Jiang, J.; Ma, J.; Yang, Y.; Liu, W.; Liu, Y. Degradation of atrazine by UV/chlorine: Efficiency, influencing factors, and products. *Water Res.* 2016, 90, 15–23.

16. Guo, K.; Wu, Z.; Shang, C.; Yao, B.; Hou, S.; Yang, X.; Song, W.; Fang, J. Radical chemistry and structural relationships of PPCP degradation by UV/chlorine treatment in simulated drinking water. *Environ. Sci. Technol.* 2017, 51, 10431–10439.

17. Wu, Z.; Fang, J.; Xiang, Y.; Shang, C.; Li, X.; Meng, F.; Yang, X. Roles of reactive chlorine species in trimethoprim degradation in the UV/chlorine process: Kinetics and transformation pathways. *Water Res.* 2016, 104, 272–282.

18. He, G.; Li, C.; Dong, F.; Zhang, T.; Chen, L.; Cizmas, L.; Sharma, V.K. Chloramines in a pilot-scale water distribution system: Transformation of 17 β -estradiol and formation of disinfection byproducts. *Water Res.* 2016, 106, 41–50.

19. Sun, P.; Meng, T.; Wang, Z.; Zhang, R.; Yao, H.; Yang, Y.; Zhao, L. Degradation of organic micropollutants in UV/NH₂Cl advanced oxidation process. *Environ. Sci. Technol.* 2019, 53, 9024–9033.

20. Domingues, E.; Fernandes, E.; Gomes, J.; Martins, R.C. Advanced oxidation processes perspective regarding swine wastewater treatment. *Sci. Total Environ.* 2021, 776, 145958.

21. Deng, Y.; Zhao, R. Advanced oxidation processes (AOPs) in wastewater treatment. *Curr. Pollut. Rep.* 2015, 1, 167–176.

22. Xu, M.; Wu, C.; Zhou, Y. Advancements in the Fenton Process for Wastewater Treatment, Advanced Oxidation Processes—Applications, Trends, and Prospects, Ciro Bustillo-Lecompte; IntechOpen: London, UK, 2020.

23. Ameta, R.; Chohadia, A.K.; Jain, A.; Punjabi, P.B. Fenton and Photo-Fenton Processes. In Advanced Oxidation Processes for Wastewater Treatment: Emerging Green Chemical Technology; Elsevier: Amsterdam, The Netherlands, 2018; pp. 49–87.

24. Rekhate, C.V.; Srivastava, J.K. Recent advances in ozone-based advanced oxidation processes for treatment of wastewater- A review. *Chem. Eng. J. Adv.* 2020, 3, 100031.

25. Ikehata, K.; Li, Y. Ozone-Based Processes. In Advanced Oxidation Processes for Wastewater Treatment: Emerging Green Chemical Technology; Elsevier Inc.: Amsterdam, The Netherlands, 2018; pp. 115–134.
26. Nawrocki, J.; Kasprzyk-Hordern, B. The efficiency and mechanisms of catalytic ozonation. *Appl. Catal. B Environ.* 2010, 99, 27–42.
27. Khan, S.; Sayed, M.; Sohail, M.; Shah, L.A.; Raja, M.A. Chapter 6—Advanced Oxidation and Reduction Processes. In Advances in Water Purification Techniques: Meeting the Needs of Developed and Developing Countries; Elsevier: Amsterdam, The Netherlands, 2018.
28. Ameta, R.; Solanki, M.S.; Benjamin, S.; Ameta, S.C. Photocatalysis. In Advanced Oxidation Processes for Wastewater Treatment: Emerging Green Chemical Technology; Elsevier Inc.: Amsterdam, The Netherlands, 2018; pp. 135–175.
29. Sonawane, G.H.; Patil, S.P.; Sonawane, S.H. Nanocomposites and Its Applications. In Applications of Nanomaterials. Advances and Key Technologies, Micro and Nano Technologies; Elsevier: Amsterdam, The Netherlands, 2018; pp. 1–22.

Retrieved from <https://encyclopedia.pub/entry/history/show/32473>